

## Esters

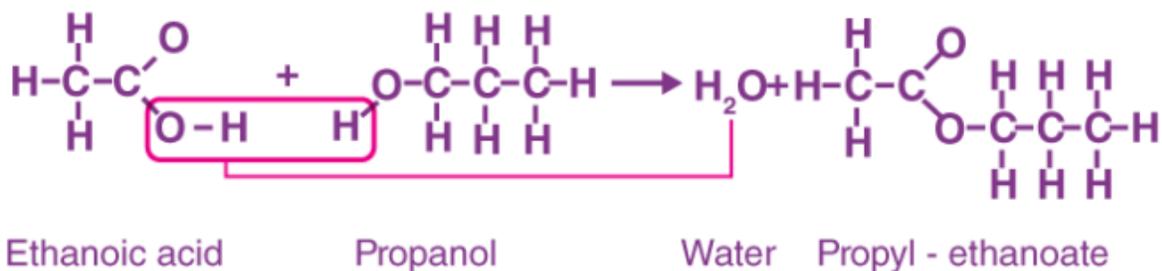
Carboxylic acids react with alcohols to form Esters and process is called as esterification.

### Drawing and naming esters

- Write the structural formula of the carboxylic acid first
- Write the formula of the alcohol beside the acid with OH on the left side

Conditions for making esters

Heat under reflux



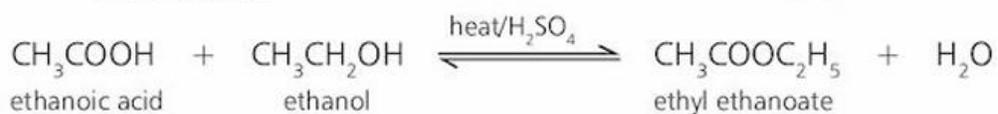
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Sulfuric acid is used as a catalyst in esterification.

Isomers of Esters: Esters have isomers example methyl butanoate and butyl methanoate are isomers of each other.

In this reaction sulfuric acid acts as a catalyst. The  $\text{—}\overset{\text{O}}{\parallel}{\text{C}}\text{—O—}$  group formed is called an ester linkage. The water given off comes partly from the acid and partly from the alcohol:



Esters are named after the acid from which they are made. So the *ethanoate* comes last and the name of the alkyl group of the alcohol, in this case *ethyl*, comes first.

The naming of esters is based on the name of the carboxylic acid and alcohol used to make them.

- The name begins with the alkyl group from the alcohol.
- The name ends with the part coming from carboxylic acid, but *–oic acid* is changed to *–oate*.

