

How to remember the Reactivity Series?

P lease	P otassium	
S top	S odium	
C alling	C alcium	
M e	M agnesium	
A	A luminium	
C areless	(C arbon)	
Z ebra	Z inc	
I nstead	I ron	
T ry	T in	
L earning	L ead	
H ow	(H ydrogen)	
C opper	C opper	
S aves	S ilver	
G old	G old	

Most reactive

Least reactive

Reactivity Series

Metal (Ion)	Reactivity
Potassium (K ⁺)	Reacts with water
Sodium (Na ⁺)	
Lithium (Li ⁺)	
Barium (Ba ²⁺)	
Strontium (Sr ²⁺)	
Calcium (Ca ²⁺)	
Magnesium (Mg ²⁺)	Reacts with acid
Aluminum (Al ³⁺)	
Manganese (Mn ⁺)	
Zinc (Zn ²⁺)	
Chromium (Cr ²⁺)	
Iron (Fe ²⁺)	
Cadmium (Cd ²⁺)	
Cobalt (Co ²⁺)	
Nickle (Ni ²⁺)	
Tin (Sn ²⁺)	
Lead (Pb ²⁺)	
Hydrogen (H ⁺)	Included for comparison
Antimony (Sb ²⁺)	Highly unreactive
Bismuth (Bi ²⁺)	
Copper (Cu ²⁺)	
Mercury (Hg ²⁺)	
Silver (Ag ⁺)	
Gold (Au ³⁺)	
Platinum (Pt ⁺)	

Most reactive



Least reactive

Element	Reaction with dilute hydrochloric acid
potassium	very violent- very explosive
sodium	very violent- explosive
calcium	very rapid- lots of hydrogen produced
magnesium	rapid- bubbles of hydrogen produced steadily
zinc	slow- bubbles of hydrogen produced slowly
iron	slow reaction- some bubbles produced
hydrogen	no reaction
copper	no reaction



